

## pH calculations

Solution	[H <sub>3</sub> O <sup>+</sup> ]	[OH <sup>-</sup> ]	pH
$9.56 \times 10^{-5} \text{ mol L}^{-1}$ of OH <sup>-</sup> ions			
0.133 mol L <sup>-1</sup> solution of HCl			
NaOH with a pH of 12.8.			
$2.47 \times 10^{-2} \text{ mol L}^{-1}$ HCl solution			
0.0534 mol L <sup>-1</sup> NaOH solution			
0.0498 mol L <sup>-1</sup> hydrochloric acid			
0.251 mol L <sup>-1</sup> sodium hydroxide			
			a sample of polluted rainwater is 4.62.
0.108 mol L <sup>-1</sup> hydrochloric acid			
			hydrochloric acid with a pH of 1.58.
0.362 mol L <sup>-1</sup> sodium hydroxide.			
0.0376 mol L <sup>-1</sup> HCl solution			
$2.48 \times 10^{-4} \text{ mol L}^{-1}$ NaOH solution			
0.00112 mol L <sup>-1</sup> HCl solution.			
$3.68 \times 10^{-2} \text{ mol L}^{-1}$ NaOH solution.			
0.125 mol L <sup>-1</sup> HCl.			
			NaOH solution at pH 10.2.
0.124 mol L <sup>-1</sup> NaOH.			
0.1 mol L <sup>-1</sup> HCl			
0.01 mol L <sup>-1</sup> HCl			
0.1 mol L <sup>-1</sup> NaOH			
	0.0350		
			10.8
		$5.66 \times 10^{-6}$	